PAPER SOLUTION

From Meerut

JEE MAIN 2026

JAN	SHIFT
28	1st

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JEE MAIN 2026 LIVE PAPER DISCUSSION

#Q. In Carius method of estimation of 'Br', 1.53 g of an organic compound gave 1 g AgBr. The % of Br in organic compound is, (Atomic mass of Ag, Br = 108, 80 u respectively)

A 35.23

B 43.53

C 27.81

D 22.71

Ans. (C)

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#Q. In period 4 of the periodic table which elements have the highest and lowest atomic radii respectively.

A. K and Br

B. Na and Cl

C. K and Se

D. Rb and Br

Ans. (A)

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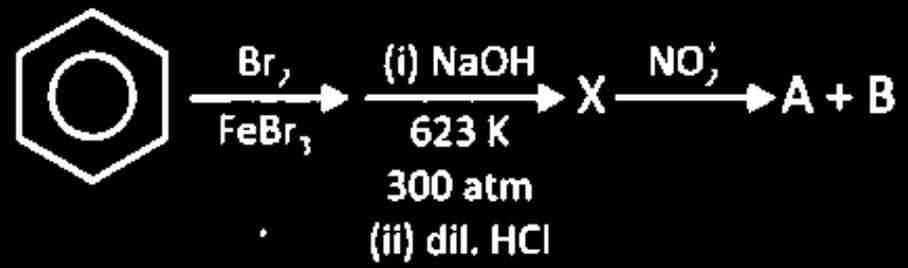
#Q. Consider the following reaction sequence :



The organic product 'A' and 'B' can be separated by

(A) Steam distillation

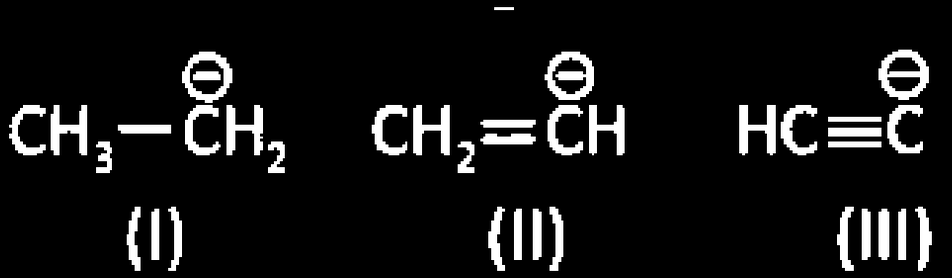
(B) Fractional distillation

(C) Distillation under reduced pressure

(D) Azeotropic distillation

Ans. (A)

#Q. Consider following ions



Stability of ions is in order.

A. III > II > I

B. II > III > I

C. I > II > III

D. I > III > II

Ans. (A)

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#Q. For a first order reaction, X → Y + Z, time required for decomposition of $\frac{1}{8}th$ and $\frac{1}{10}th$ of its initial conc. is $t_{1/8}$ and $t_{1/10}$.

the value of $\left(t_{1/8}/t_{1/10}\right) \times 10$ =

Take : log 8 = 0.90, log 7 = 0.84, log 9 = 0.95

A 9

B 10

C 12

D 8

Ans. (C)

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#Q. Choose the correct statements in respect of hydrides of Group-15.

A. Reducing power increasing down the group.

B. Basic nature increases down the group.

C. Stability decreases down the group.

D. Boiling point decreases regularly down the group.

(A) A, B and C only

(B) A, B and D only

(C) A and C only

(D) B, C and D only

Ans. (C)

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#Q. Which is correct option.

(A) $[Ni(CN)_4]^{2-}$ is paramagnetic while $[NiCl_4]^{2-}$ and $[Ni(CO)_4]$ are diamagnetic.

(B) $[Ni(CN)_4]^{2-}$ and $[Ni(CO)_4]$ are diamagnetic while $[NiCl_4]^{2-}$ is paramagnetic.

(C) $[NiCl_4]^{2-}$ and $[Ni(CO)_4]$ are paramagnetic while $[Ni(CN)_4]^{2-}$ is diamagnetic.

(D) $[NiCl_4]^{2-}$ and $[Ni(CN)_4]^{2-}$ are paramagnetic while $[Ni(CO)_4]$ is diamagnetic.

A A

B B

C C

D D

Ans. (B)

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#Q. The wave number of three spectral lines of H-atom are given. Identify the correct set of spectral lines belonging to Balmer series

A. 5R/36, 3R/16, 21R/100

B. 3R/4, 3R/16, 7R/144

C. 7R/144, 3R/16, 16R/255

D. 5R/36, 3R/16, 21R/24

Ans. (A)

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#Q. Given below are two statements.

Statement I : Among XeF_4, BF_4^- and SF_4 the species having equal M-X bond lengths are XeF_4 and BF^-_4.

(M = central atom).

Statement II : Among O_2^{2-}, O_2^-, F_2 and O_2^+ the highest bond order is for F_2 and O_2^{2-}.

In the light of the above statements, choose the most appropriate option.

(A) Both statement-I and statement-II are correct

(B) Both statement-I and statement-II are incorrect

(C) Statement-I is correct but statement-II is incorrect

(D) Statement-I is incorrect but statement-II is correct

Ans. (C)

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#Q. Among the following coloured ion is/are:

(A) Ti^{3+} and V^{3+}

(B) Ti^{3+} and Sc^{3+}

(C) Ti^{4+} and V^{3+}

(D) V^{2+} and Sc^{3+}

Ans. (A)

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#Q. At T(K), 2 moles of liquid A and 3 moles of liquid B are mixed. The vapour pressure of ideal solution SO formed is 320 mm Hg. At this stage one mole of A are mixed further, the vapour pressure is found to be 340 mm Hg. The vapour pressure of pure A and B are respectively

A 200 mm Hg, 400 mm Hg

B 440 mm Hg, 240 mm Hg

C 300 mm Hg, 400 mm Hg

D 240 mm Hg, 440 mm Hg

Ans. (B)

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#Q. Observe the following reaction:

$$C_3H_6Cl_2 \xrightarrow[NaNH_2]{excess} (X) \xrightarrow[H_2O]{Hg^{2+},H^+} (Y)$$

The product (Y) gives which of the following test?

A. Tollen's test

B. Lucas test

C. Iodoform test

D. Fehling's test

Ans. (C)

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#Q. $X \xrightarrow[H_3O^+]{(i)CO_2,500atm} Y$

X react with $FeCl_3$
X contain C = 76.57% H = 6.43 % O = 17%
V.D. of X = 47
incorrect statement among following.

(A) X reacts with $NaHCO_3$

(B) X is more acidic than Y

(C) Y is salicylic acid

(D) Y is product of Kolbe's reaction

Ans. (B)

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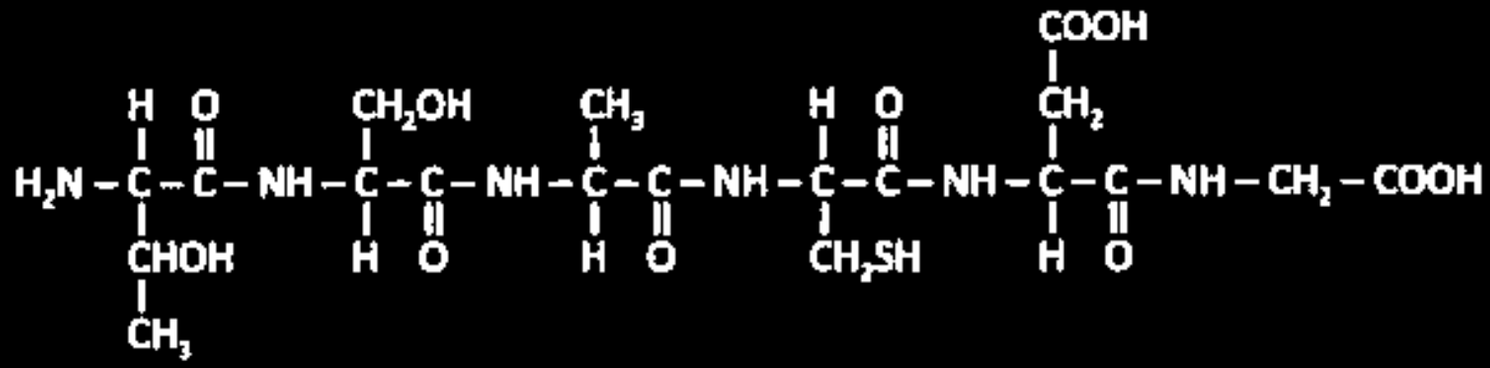
#Q. **Consider the following polypeptide :**
In the given polypeptide, Y is the essential amino acid present. The correct representation of Y and the name of amino acids in the correct sequence in polypeptide is



A Polypeptide (name of amino acid), Thr-Ser-Ala-Cys-Asp-Gly

B Polypeptide (name of amino acid), Ser-Ala-Thr-Cys-Asp-Gly

C Polypeptide (name of amino acid), Thr-Ser-Cys-Asp-Ala-Gly

D Polypeptide (name of amino acid), Thr-Ser-Ala-Asp-Cys-Gly

Ans. (A)

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#Q. Which of the following reaction is correctly matched with the product formed?

A 

B 

C 

D 

Ans. (A)

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#Q. Match the column-I showing compounds with column-II showing suitable test for that compound.

	Column-I		Column-II
(P)	$C_6H_5COCH_2CH_3$	a	Iodoform test
(Q)	C_6H_5CHO	b	2, 4-DNP test
(R)	$C_6H_5CH_2CHO$	c	Tollen test
(S)	$C_6H_5COCH_3$	d	Fehling test

	P	Q	R	S
A	b	b, c	b, c, d	a, b
B	b	b, c, d	b, c, d	a, b, c
C	a, b	b, c, d	b, c, d	a, b, d
D	b	b, c, d	b, c	a, b

Ans. (A)

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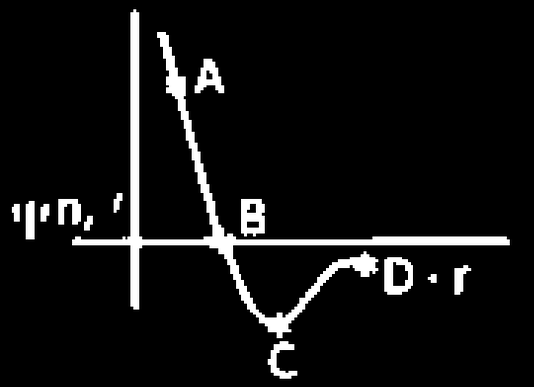
#Q. Consider the diagram



Radical node is shown by

(A) A

(B) B

(C) C

(D) D

Ans. (B)

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#Q. In reversible isothermal process at 600 K, pressure changes from 0.5 MPa to 0.2 Mpa, then find ΔU, W and q. Given moles of gas in container is 1 mol. (R = 8.3 JK^{-1} mol^{-1}).

A
$\Delta U = 0$
$q = -4.587$ kJ
$w = +4.587$ kJ

B
$\Delta U = 0$
$q = 0$
$w = 0$

C
$\Delta U = 0$
$q = 0$
$w = -4.587$ kJ

D
$\Delta U = 0$
$q = +4.587$ kJ
$w = -4.587$

Ans. (D)

#Q. Calculate the sum of number of geometrical isomers of $[MCl\ BrNO_2CN]$, number of optically inactive isomers of $[M(OX)_2Cl_2]$ and number of geometrical isomers of $[MCl_3Br_3]$

Ans. (15)

#Q. Consider a galvanic cell reaction :

$$3BH_4^-(aq) + 4ClO_3^-(aq) \xrightarrow{H^+} 4Cl^- + 3H_2BO_3^-(aq) + 3H_2O$$

EMF of cell is given by,

$$E_{cell} = E^o_{cell} - \frac{RT}{nF}\ln[Q]$$

Here 'Q' is reaction quotient for the given cell reaction. Find 'n'.

Ans. (24)

#Q. 500 ml of 1.2 M KI solution is reacting with 500 ml of 0.2 M $KMnO_4$ solution, and product iodine is further reacting with 0.1 M $Na_2S_2O_3$ solution. The volume of $Na_2S_2O_3$ solution required for complete reaction is_____ ml.

Ans. (5000)

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#Q. For equivalence point X ml of 0.02 M HCl is treated with 5 mL of 0.02 M of a weak base. The pK_b of weak base is 5.69 and the pH of the resulting solution is Y at half of the equivalence point. The value of (x + y) is:

A 5

B 8.81

C 13.31

D 3.81

Ans. (C)